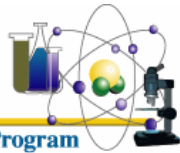




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Precipitation and Deposition of Aluminum-Containing Phases in Tank Wastes

Shas V. Mattigod, Andrew Felmy, Li Wang

Pacific Northwest National Laboratory, Richland, Washington

David T. Hobbs

Savannah River National Laboratory, Aiken, South Carolina

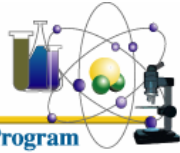
Ilhan Aksay , Dan Dabbs

Princeton University, Princeton, New Jersey

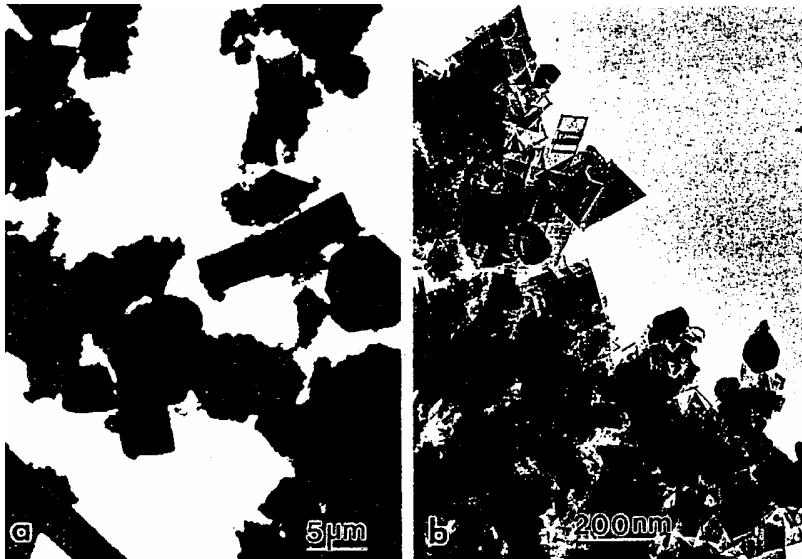


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Relevance

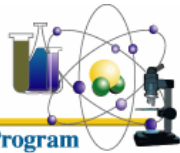


- ✓ Tank sludge retrieval and pretreatment precede vitrification.
- ✓ Al is a major component of both sludge and supernate fractions of HLW.
- ✓ Minimization of HLW glass volume requires sludge washing and leaching to dissolve Al-containing wastes.



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Processing Hiatus



∨ Scaling and clogging from Al-Si phases:

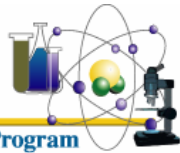
- ∅ 2H evaporator at SRS shutdown
- ∅ Plugging of Cs-removal columns at SRS from mineral formation.
- ∅ Occasional blocked pipes at Hanford tank farm due to aluminous precipitates.

∨ Processing hiatus results in escalated cost and extended time for treating tank wastes

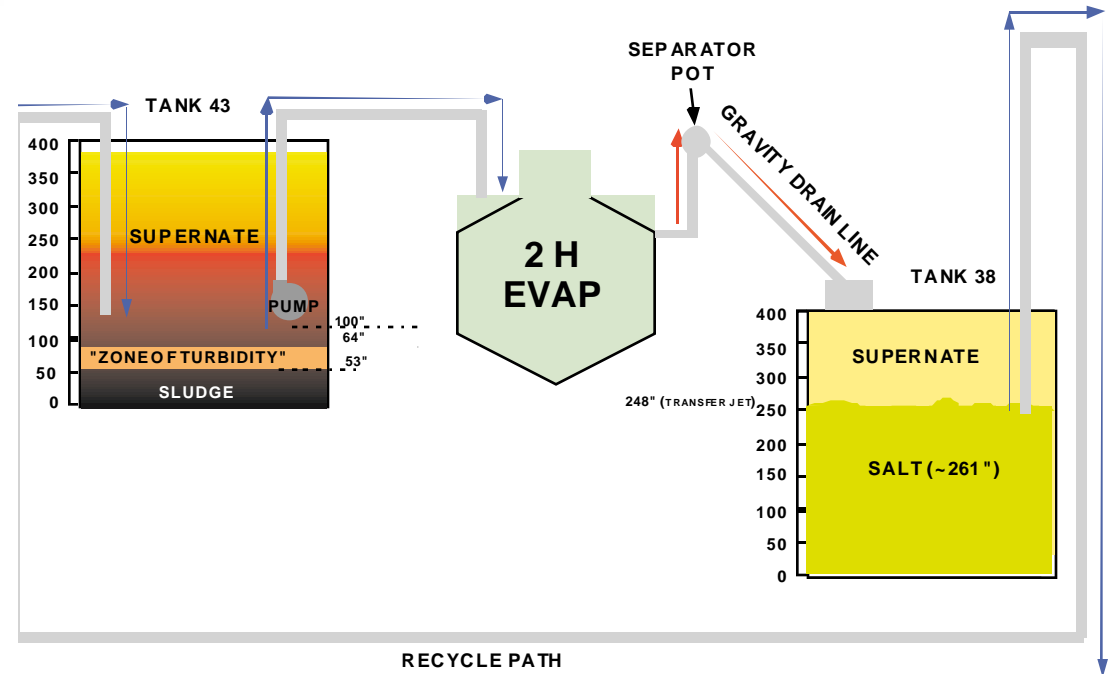


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SRS 2H-Evaporator



✓ Precipitates

Plugged concentrate line (97 - 98)
>3000 kg solids in evaporator (99)

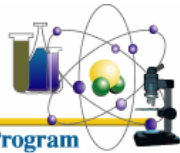
✓ Down time and Cost

Gravity Drain Line: 4 mo - \$4M
Evaporator: >22 mo, \$10+ M



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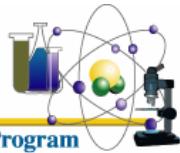
Critical Need

- ∨ Limited knowledge about mechanisms of formation and transformation of Al-Si phases under tank and pretreatment conditions (Si/Al ~ 0.003, high salt and OH⁻, range of temperature)
- ∨ Understand factors that control the extent and the rate of formation of Al-Si phases that form hard cementitious scales
- ∨ Develop process schemes to avoid/inhibit formation of cementitious Al-Si phases.



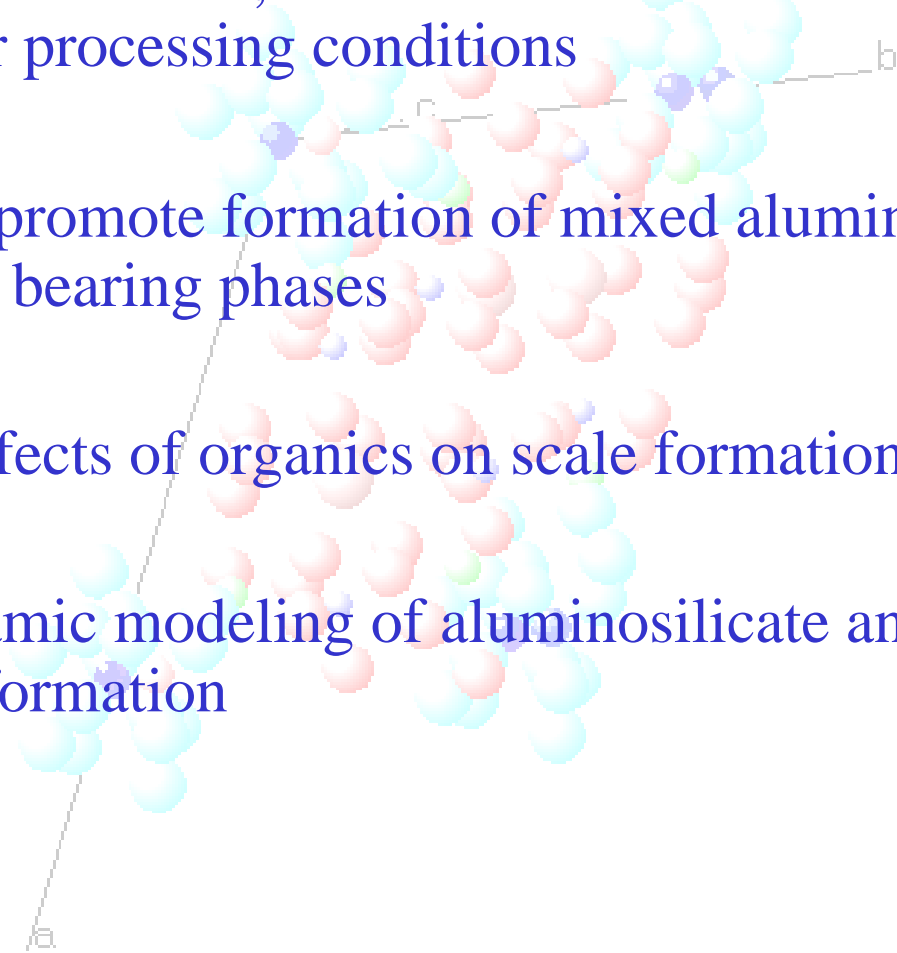
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Objectives

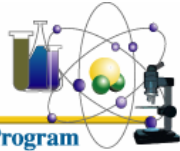
- ✓ Formation, solubilities, and transformation of Al-bearing phases under processing conditions
- ✓ Factors that promote formation of mixed aluminosilicate and uranium bearing phases
- ✓ Inhibiting effects of organics on scale formation.
- ✓ Thermodynamic modeling of aluminosilicate and uranium solid phase formation





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Previous Work

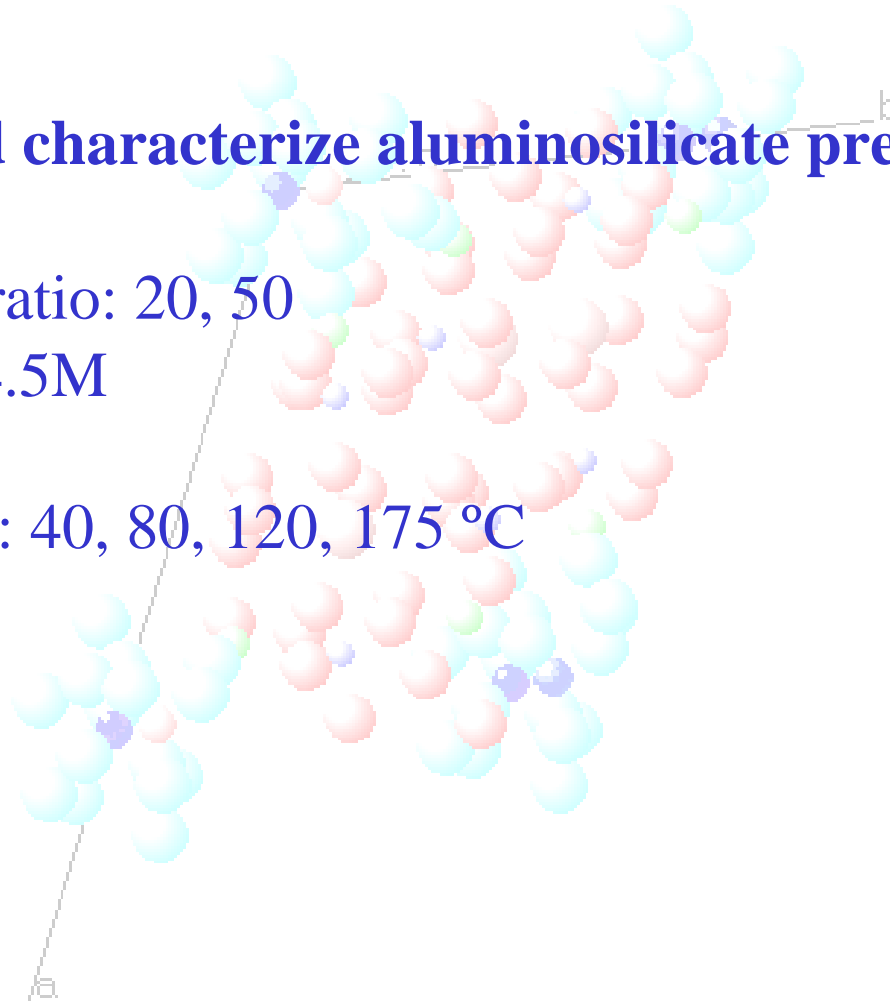
Identify and characterize aluminosilicate precipitates

Al/Si molar ratio: 20, 50

OH: 0.1, 1, 4.5M

NaNO₃: 3M

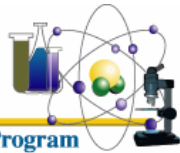
Temperature: 40, 80, 120, 175 °C





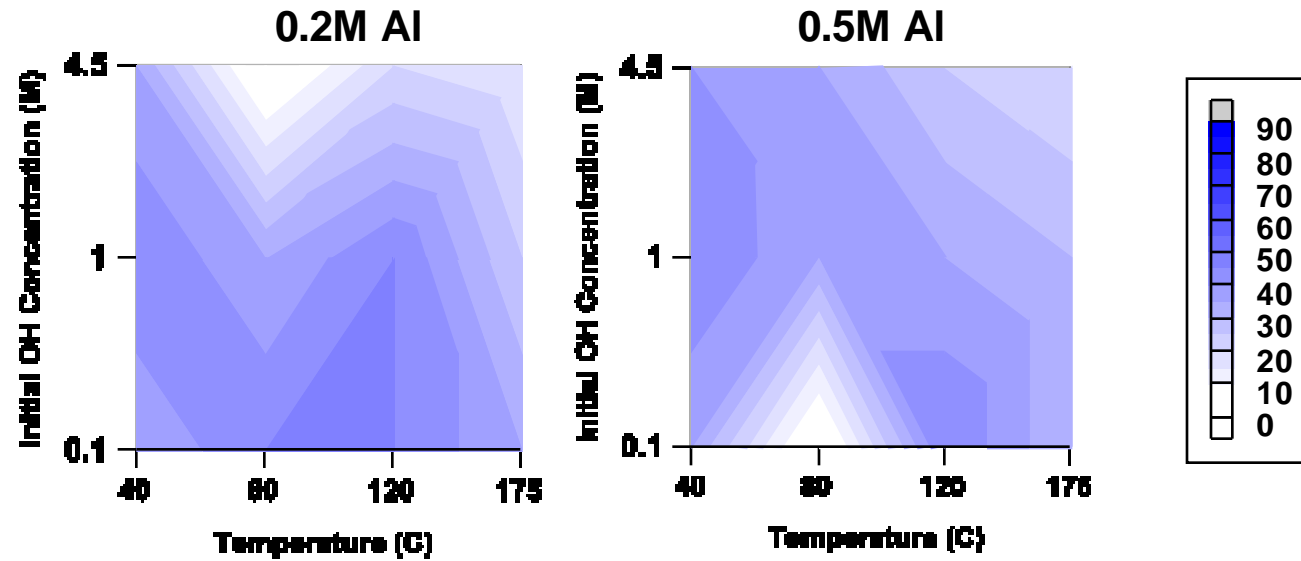
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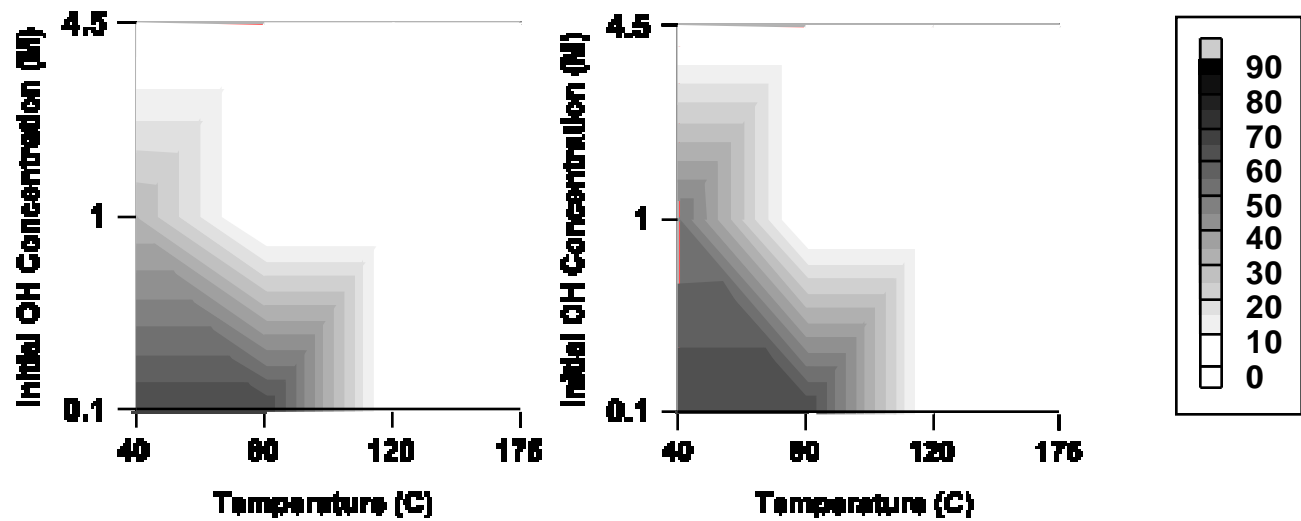


Predominance Diagrams

Precursor Phase



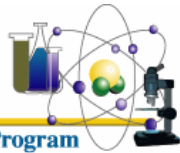
Zeolite A





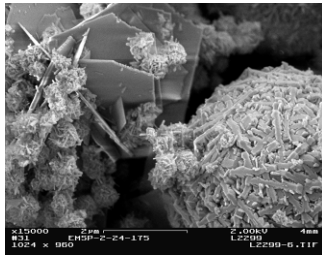
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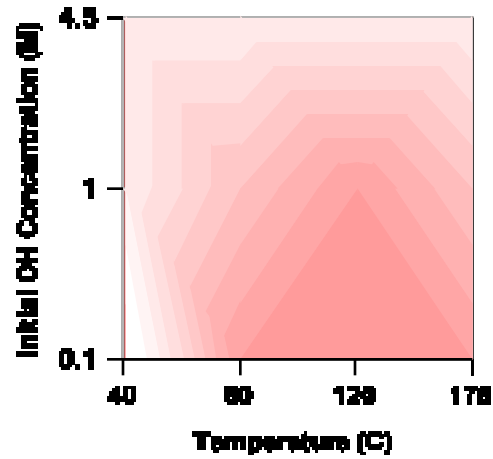


Predominance Diagrams

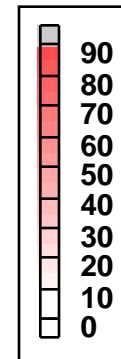
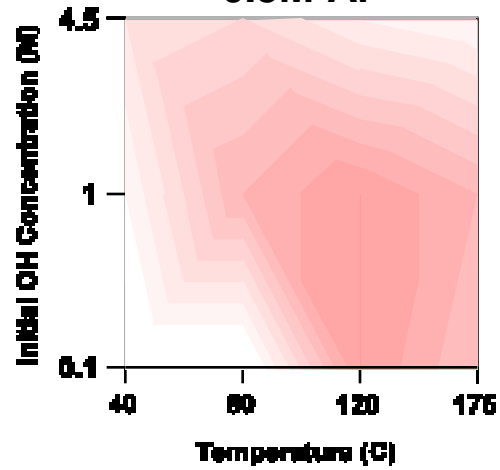
Sodalite



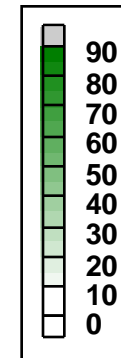
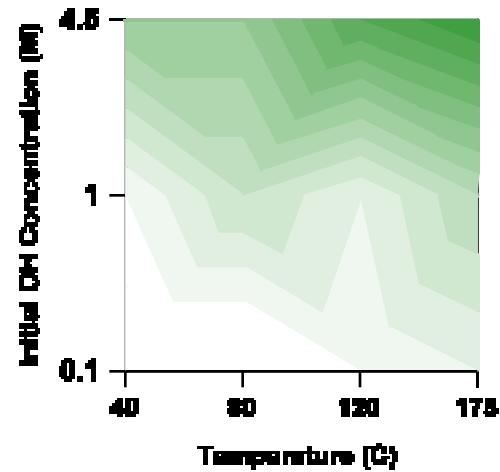
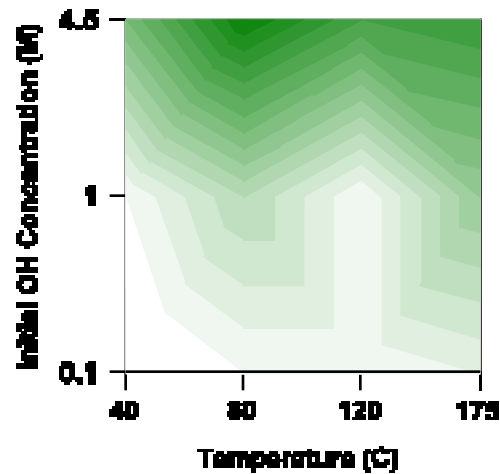
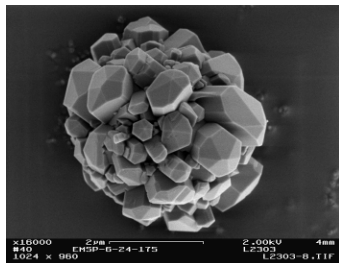
0.2M Al



0.5M Al



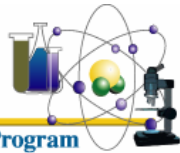
Cancrinite



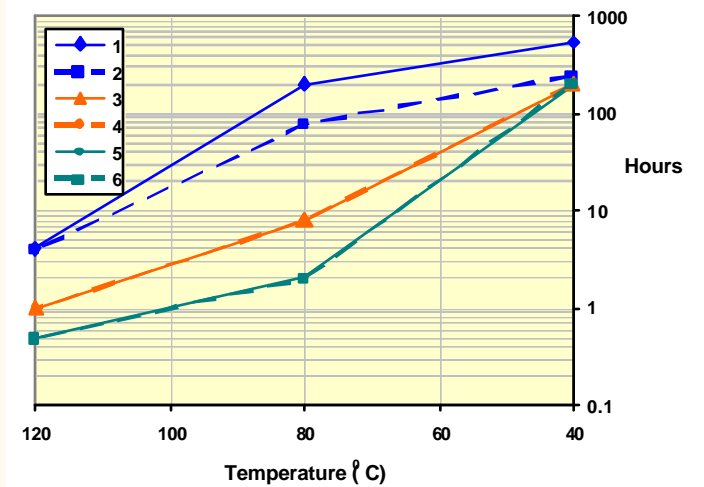
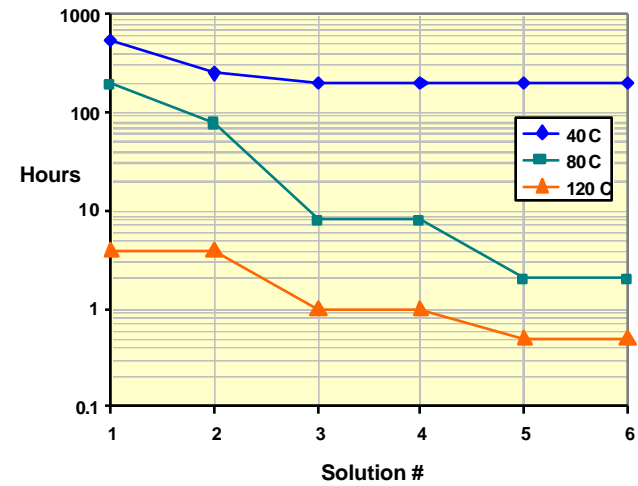
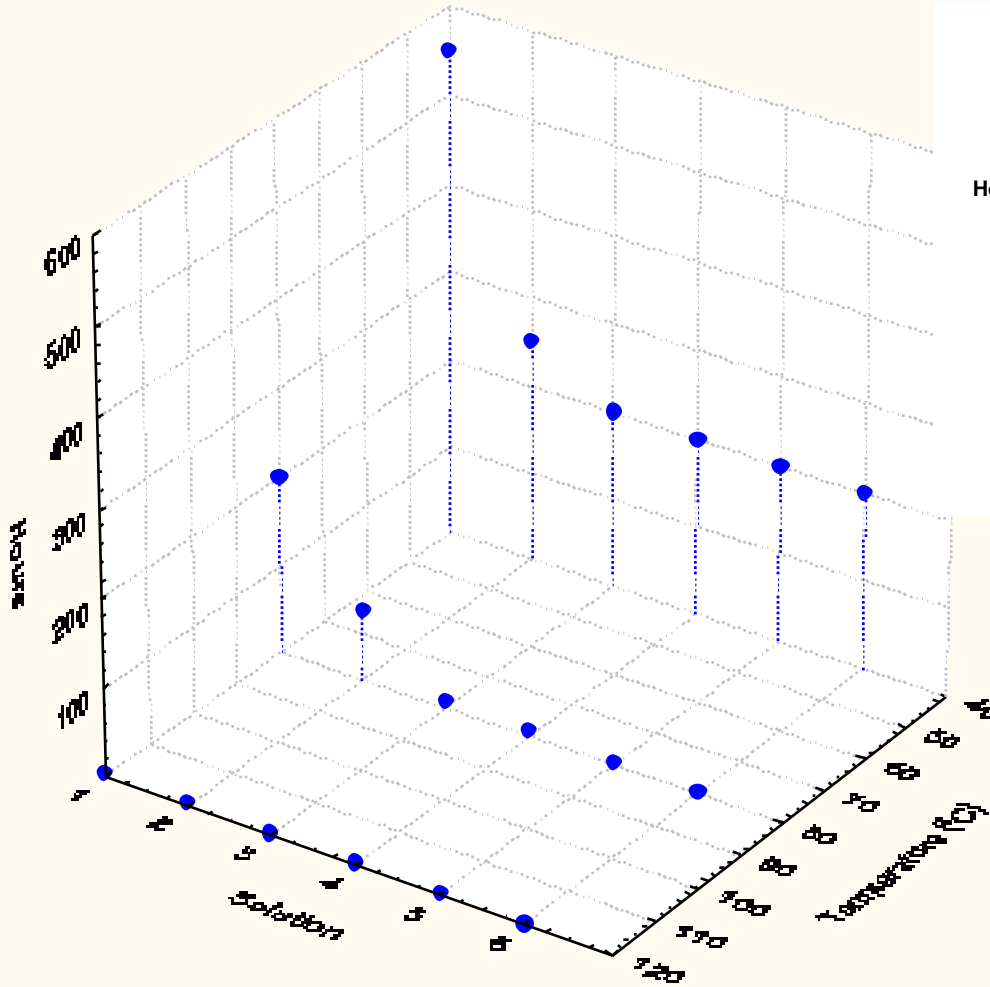


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Crystallization Kinetics





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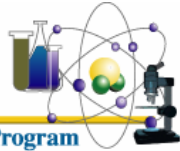
Synopsis

- ***Solid Phase Formation***
Amorphous precursor phase precipitates initially - with time converts to mainly zeolitic crystalline phases
- ***Al/Si Ratio in Precipitates***
Precipitates become more aluminous with increasing temperature and Al concentration
- ***Crystallization Kinetics***
Higher OH⁻, increasing temperature promote more rapid crystallization of the precursor phase
- **Dominant Crystalline Phases**
<80 C - <1M OH⁻: zeolite A, sodalite
>80 C - <1M OH⁻: sodalite, cancrinite
>80 C - >1M OH⁻ cancrinite, sodalite



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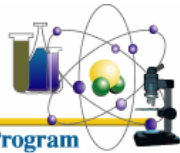
Summary

- ✓ Provide sound scientific knowledge for processing schemes to avoid and/or inhibit formation and growth of Al-Si phases
- ✓ Knowledge of aluminosilicate chemistry critical to glass waste minimization
- ✓ Develop insights into industrial fouling problems
- ✓ Gather data on geochemistry of two most abundant and ubiquitous elements (Al and Si) in earth's crust



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Research Plan (2005 – 2007)

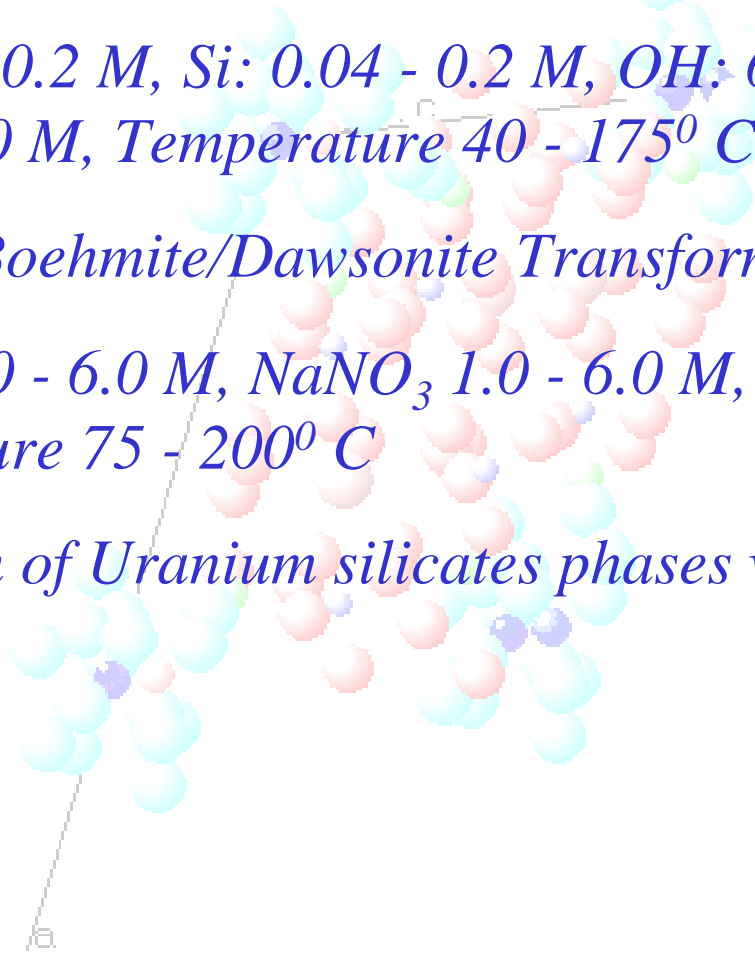
∨ Formation and solubilities of aluminosilicates

*Al: 0.01 – 0.2 M, Si: 0.04 - 0.2 M, OH: 6.0 – 10.0 M,
NaNO₃ 5.0 M, Temperature 40 - 175^o C*

∨ *Gibbsite/Boehmite/Dawsonite Transformation*

*NaOH: 1.0 - 6.0 M, NaNO₃ 1.0 - 6.0 M, NaNO₂: 1.0 – 3.0M
Temperature 75 - 200^o C*

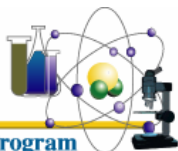
∨ *Formation of Uranium silicates phases with NAS*





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Research Plan (2005 – 2007)

✓ Role of organics in inhibiting precipitation and scale formation

- 1. Low-chain polyols to stabilize aluminosilicate particles.*
- 2. Polyelectrolytes and diblock copolymers to prevent nucleation and particle growth*

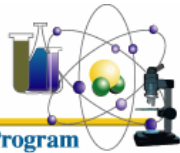
✓ Thermodynamic Modeling

- 1. ^{27}Al and ^{29}Si NMR – determine speciation of Al and Si under relevant conditions for use in Pitzer model development*
- 2. Model U(VI) solution chemistry in high ionic strength solutions*



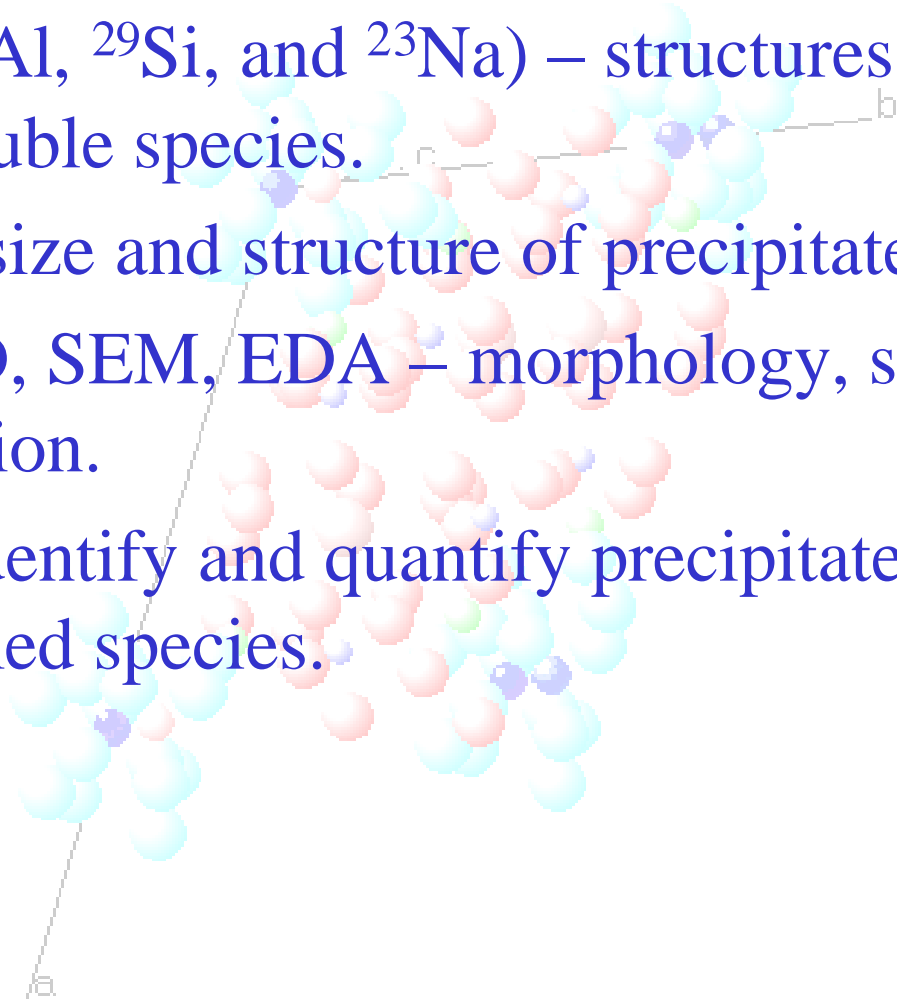
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Approach

- ✓ NMR (^{27}Al , ^{29}Si , and ^{23}Na) – structures of soluble and insoluble species.
- ✓ SAXS – size and structure of precipitates.
- ✓ TEM, ED, SEM, EDA – morphology, structure and composition.
- ✓ XRD – identify and quantify precipitated and transformed species.





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Integration

